

## CHEMISTRY A LEVEL PAPER 1 MARK SCHEME

Question number	Answer	Additional guidance	Marks
<b>1(a)</b>	B		<b>1</b>
<b>1(b)</b>	B		<b>1</b>
<b>1(c)</b>	C		<b>1</b>
<b>1(d)(i)</b>	<p>An explanation that makes reference to the following points:</p> <ul style="list-style-type: none"> <li>• the shape is V-shaped/bent (1)</li> <li>• because the three electron pairs on Sn repel one another (1)</li> </ul>	<p>First marking point can be awarded for a correctly-drawn diagram</p>	<b>2</b>
<b>1(d)(ii)</b>	<ul style="list-style-type: none"> <li>• prediction: any value between 95 and 119° (1)</li> <li>• justification: lone pair – bond pair repulsion is greater than bond pair – bond pair repulsion (and hence the angle is less than 120°) (1)</li> </ul>		<b>2</b>

**(Total for Question 1 = 7 marks)**

Question number	Answer	Additional guidance	Marks
2(a)	D		1
2(b)	B		1
2(c)	D		1
2(d)	<ul style="list-style-type: none"> <li>order: GZ &gt; DX &gt; EY (1)</li> </ul> Justification: <ul style="list-style-type: none"> <li>the ions in GZ have higher charges (than those in both EY and DX) (1)</li> <li>the ions in DX are smaller than those in EY (1)</li> </ul>		3
2(e)	<ul style="list-style-type: none"> <li>construction of balanced cycle (1)</li> <li>substitution and evaluation of 2<sup>nd</sup> IE (1)</li> </ul>	<u>Example calculation</u> $- 2258 = -590 - 2^{\text{nd}} \text{ IE} + 2 (349) - 178 - 2 (122) - 796$ hence 2 <sup>nd</sup> IE = (+) 1148 ( kJ mol <sup>-1</sup> ) correct answer, no working scores 2 marks	2

**(Total for Question 2 = 8 marks)**

Question number	Answer	Additional guidance	Marks
<b>3(a)</b>	<ul style="list-style-type: none"> <li>(V) <math>(1s^2 2s^2 2p^6 3s^2) 3p^6 3d^3 4s^2</math> (1)</li> <li>(V<sup>3+</sup>) <math>(1s^2 2s^2 2p^6 3s^2) 3p^6 3d^2</math> (1)</li> </ul>	Allow $4s^2 3d^3$	<b>2</b>
<b>3(b)(i)</b>	<p>An explanation that makes reference to the following points:</p> <ul style="list-style-type: none"> <li><math>E^\ominus</math> of redox system 5 is more negative / less positive than that of both redox systems 4 and 3 (but not than that of redox systems 2 and 1) (1)</li> <li>therefore SO<sub>2</sub> releases electrons to / reduces VO<sub>2</sub><sup>+</sup> to VO<sup>2+</sup> and then VO<sup>2+</sup> to V<sup>3+</sup> (1)</li> <li>(yellow to) blue <u>then</u> green (1)</li> </ul>	<p>Accept explanations based on calculating <math>E^\ominus_{\text{cell}}</math> for the reactions</p> <p>Accept correct use of anticlockwise rule</p> <p>Ignore green colour before blue</p>	<b>3</b>
<b>3(b)(ii)</b>	<p>The reaction is <b>not</b> feasible because:</p> <ul style="list-style-type: none"> <li><math>E^\ominus</math> of redox system 2 is less negative than that of redox system 1 (1)</li> <li>therefore V<sup>2+</sup> ions in system 2 will not release electrons to the V<sup>2+</sup> ions in system 1 (1)</li> </ul>	<p>Accept explanations based on calculating <math>E^\ominus_{\text{cell}}</math> for the reactions</p> <p>Accept correct use of anticlockwise rule</p>	<b>2</b>

**(Total for Question 3 = 7 marks)**

Question number	Answer	Additional Guidance	Marks
<b>4(a)</b>	<ul style="list-style-type: none"> <li>axes: correct way round, labelled, suitable scale (1)</li> <li>all points plotted correctly to nearest gridline AND straight line through first 3 points, straight line through last 3 points, and the two lines intersect (1)</li> </ul>	The scale is suitable if the distance between the first point plotted and the last point plotted covers more than half of the graph paper on each axis.	<b>2</b>
<b>4(b)(i)</b>	10 (cm <sup>3</sup> ) (1)	Award mark for a value read correctly from the candidate's graph	<b>1</b>
<b>4(b)(ii)</b>	(0.005 × 0.5) = 2.5(0) × 10 <sup>-3</sup> / 0.0025 (mol) (1)		<b>1</b>
<b>4(b)(iii)</b>	(0.010 × 1.00) = 1 × 10 <sup>-2</sup> / 0.01 (mol) (1)	Answer to (iii) csq on (ii)	<b>1</b>
<b>4(b)(iv)</b>	x = 4 and y = 2 (1)	Answer to (iv) csq on (i) and (ii), but x + y must total 6	<b>1</b>
<b>4(c)</b>	$\text{Cu}^{2+}(\text{aq}) + 2\text{OH}^{-}(\text{aq}) \rightarrow \text{Cu}(\text{OH})_2(\text{s})$ <ul style="list-style-type: none"> <li>correctly balanced equation (1)</li> <li>state symbols (1)</li> </ul>	Allow $[\text{Cu}(\text{H}_2\text{O})_6]^{2+}(\text{aq}) + 2\text{OH}^{-}(\text{aq}) \rightarrow \text{Cu}(\text{H}_2\text{O})_4(\text{OH})_2(\text{s}) + 2\text{H}_2\text{O}(\text{l})$	<b>2</b>

**(Total for Question 4 = 8 marks)**

Question number	Answer	Additional guidance	Marks
<b>5(a)</b>	$\text{Cl}_2 + 2\text{NaOH} \rightarrow \text{NaOCl} + \text{NaCl} + \text{H}_2\text{O}$ (1)	Accept $\text{Cl}_2 + 2\text{OH}^- \rightarrow \text{OCl}^- + \text{Cl}^- + \text{H}_2\text{O}$ Ignore state symbols	<b>1</b>
<b>5(b)</b>	An explanation that makes reference to the following points: <ul style="list-style-type: none"> <li>• Cl has oxidation number +4 to in <math>\text{ClO}_2</math> (1)</li> <li>• Cl (in <math>\text{ClO}_2</math>) changes oxidation number to +3 (<math>\text{ClO}_2^-</math>) and +5 (<math>\text{ClO}_3^-</math>) (1)</li> <li>• (therefore) chlorine/it (in <math>\text{ClO}_2</math>) has been both oxidised and reduced (1)</li> </ul>	Accept 4+ Accept 3+ and 5+ Allow answers in any order	<b>3</b>
<b>5(c)(i)</b>	A		<b>1</b>
<b>5(c)(ii)</b>	An answer that makes reference to the following points: <ul style="list-style-type: none"> <li>• Mn changes from +7 to +2 and Cl changes from -1 to 0 (1)</li> <li>• therefore ratio is <math>\text{MnO}_4^-</math> to <math>5\text{Cl}^-</math> /the ratio of <math>\text{MnO}_4^-</math> to <math>\text{Cl}^-</math> is 1 to 5 (1)</li> </ul>	Accept 7+ and 2+ Accept 1-	<b>2</b>
<b>5(d)(i)</b>	An explanation that makes reference to the following points: <ul style="list-style-type: none"> <li>• the HBr dissolves in water (in the air) (1)</li> <li>• and forms droplets (of hydrobromic acid) (1)</li> </ul>		<b>2</b>
<b>5(d)(ii)</b>	<ul style="list-style-type: none"> <li>• the (conc.) <math>\text{H}_2\text{SO}_4</math> is reduced to <math>\text{SO}_2</math> (1)</li> </ul>	Accept redox	<b>1</b>

**(Total for Question 5 = 10 marks)**

Question number	Answer	Additional guidance	Marks
<b>6(a)</b>	<ul style="list-style-type: none"> <li>• determines number of mol of CO<sub>2</sub> produced (1)</li> <li>• calculates number of mol NaO<sub>2</sub> in 880 g (1)</li> <li>• calculates the number of mol of CO<sub>2</sub> reacted with NaO<sub>2</sub> and hence determines the number of mol of CO<sub>2</sub> in excess (1)</li> <li>• calculation of mass of Na<sub>2</sub>O = 248 g (1)</li> </ul>	<p><u>Example of calculation</u></p> <p>2 mol glucose produces 12 mol CO<sub>2</sub>  880 g NaO<sub>2</sub> = 16 mol  16 mol NaO<sub>2</sub> remove 8 mol CO<sub>2</sub> (therefore 4 mol CO<sub>2</sub> is excess)  mass of Na<sub>2</sub>O required is (4 x 62) = 248 g</p> <p>3<sup>rd</sup> mark is csq on answers given in 1<sup>st</sup> and 2<sup>nd</sup> marks  4<sup>th</sup> mark is csq on answer given in 3<sup>rd</sup> mark</p> <p>Correct answer with units and no working scores 4 marks.</p>	<b>4</b>
<b>6(b)</b>	<p>An explanation that makes reference to the following points:</p> <div style="text-align: center;"> </div> <p style="text-align: right;">(1)</p> <ul style="list-style-type: none"> <li>• it is a radical because it has an unpaired electron (1)</li> </ul>	<p>Accept any other symbols in place of dots and crosses, including a minus sign to replace the triangle</p>	<b>2</b>

Question number	Answer	Additional guidance	Marks
6(c)	<ul style="list-style-type: none"> <li>• calculation of number of moles Na<sub>2</sub>O<sub>2</sub> (= mols H<sub>2</sub>O<sub>2</sub>)</li> <li>• calculation of number of moles Ce<sup>4+</sup></li> <li>• determination of the ratio H<sub>2</sub>O<sub>2</sub> : Ce<sup>4+</sup></li> <li>• formula of cerium ion is Ce<sup>3+</sup></li> </ul>	<p><u>Example of answer</u></p> <p>(1) <math>n(\text{Na}_2\text{O}_2) = (0.403 \div 78.0) = 5.17 \times 10^{-3}</math>  <math>n(\text{Ce}^{4+}) = (24.85 \div 1000 \times 0.420) = 1.04 \times 10^{-2}</math></p> <p>(1) <math>n(\text{Ce}^{4+}) : n(\text{H}_2\text{O}_2) = 1.04 \times 10^{-2} : 5.17 \times 10^{-3} = 2 : 1</math></p> <p>(1) 1<sup>st</sup> and 2<sup>nd</sup> mark: accept answers to any significant figures, except 1s.f</p> <p>3<sup>rd</sup> mark csq on answers given in 1<sup>st</sup> and 2<sup>nd</sup> marks            4<sup>th</sup> mark csq on answer given in 3<sup>rd</sup> mark</p>	<b>4</b>

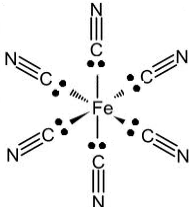
**(Total for Question 6 = 10 marks)**

Question number	Answer	Additional guidance	Marks
7(a)	B		1
7(b)(i)	<ul style="list-style-type: none"> <li>• calculation of <math>n(\text{CO})</math> at equilibrium and <math>n(\text{H}_2)</math> at equilibrium (1)</li> <li>• converting number of moles to concentration (1)</li> <li>• evaluation of <math>K_c</math> by substitution (1)</li> <li>• correct answer to 3 sf (1)</li> <li>• units: <math>\text{dm}^6 \text{mol}^{-2}</math> (1)</li> </ul>	<p><u>Example of calculation</u>  <math>n(\text{CO})</math> at equilibrium = 0.114 (mol)  <math>n(\text{H}_2)</math> at equilibrium = 0.228 (mol)  <math>[\text{CO}] = 0.0950 \text{ mol dm}^{-3}</math>, <math>[\text{H}_2] = 0.190 \text{ mol dm}^{-3}</math>,  <math>[\text{CH}_3\text{OH}] = 0.0717 \text{ mol dm}^{-3}</math>  <math>K_c = 0.0717 \div (0.0950 \times 0.190^2) = 20.9068\dots</math>  <math>= 20.9 \text{ dm}^6 \text{mol}^{-2}</math></p> <p>Allow <math>\text{mol}^{-2} \text{dm}^6</math></p> <p>3<sup>rd</sup> and 4<sup>th</sup> marks csq on answers given in 1<sup>st</sup> and 2<sup>nd</sup> marks</p> <p>Correct final answer to 3 sf with units but no working scores 5 marks</p>	5
7(b)(ii)	<p>An explanation that makes reference to the following points:</p> <ul style="list-style-type: none"> <li>• an increase in temperature shifts the equilibrium to the left (1)</li> <li>• an increase in pressure shifts the equilibrium to the right (1)</li> <li>• these changes produce opposing effects, so to predict the effect on the yield it is necessary to know the relative effects of each one (1)</li> </ul>		3

(Total for Question 7 = 9 marks)

Question number	Answer	Additional guidance	Marks
<b>8(a)</b>	A		<b>1</b>
<b>8(b)(i)</b>	An explanation that makes reference to the following points: <ul style="list-style-type: none"> <li>• stage 2 is an equilibrium reaction / only partial ionisation occurs (1)</li> <li>• therefore fewer hydrogen ions are formed (1)</li> </ul>	Accept dissociation for ionisation	<b>2</b>
<b>8(b)(ii)</b>	<ul style="list-style-type: none"> <li>• rearrangement of equation <math>\text{pH} = -\log [\text{H}^+]</math> and substitution to give final answer (1)</li> </ul>	Example calculation: $[\text{H}^+] = 10^{-\text{pH}}$ $10^{-0.97} = 0.107 \text{ (mol dm}^{-3}\text{)}$  Allow 0.11 (mol dm <sup>-3</sup> )  Correct answer with no working scores 1 mark	<b>1</b>
<b>8(c)</b>	<ul style="list-style-type: none"> <li>• rearrangement of <math>K_a</math> expression (1)</li> <li>• calculation of <math>[\text{CH}_3\text{COOH}]</math> and <math>[\text{CH}_3\text{COO}^-]</math> (1)</li> <li>• substitution, and evaluation of <math>[\text{H}^+]</math> in the buffer solution (1)</li> <li>• conversion of <math>[\text{H}^+]</math> to pH for buffer solution (1)</li> </ul>	Example of calculation : $[\text{H}^+] = K_a \frac{[\text{CH}_3\text{COOH}]}{[\text{CH}_3\text{COO}^-]}$ $[\text{CH}_3\text{COOH}] = 0.333 \text{ mol dm}^{-3} \text{ and}$ $[\text{CH}_3\text{COO}^-] = 0.167 \text{ mol dm}^{-3}$ $[\text{H}^+] = 1.74 \times 10^{-5} \times 0.333 / 0.167$ $= 3.48 \times 10^{-5} \text{ (mol dm}^{-3}\text{)}$ so $\text{pH} = -\lg 3.48 \times 10^{-5} = 4.46$  Accept answers that use forms of the Henderson-Hasselbach equation  Correct answer with no working scores 4 marks	<b>4</b>

**(Total for Question 8 = 8 marks)**

Question number	Answer	Additional guidance	Marks
9(a)(i)	A		1
9(a)(ii)	 <ul style="list-style-type: none"> <li>• 3D shape correctly shown (1)</li> <li>• electron pairs shown OR all six bonds shown clearly to the carbons (1)</li> </ul>	Charge and square brackets not required  Allow dotted lines for the wedges going backwards  Allow arrows to represent dative covalent bonds instead of electron pairs	2
9(b)	<ul style="list-style-type: none"> <li>• step 1 – ligand substitution/ligand exchange (1)</li> <li>• step 2 – redox (1)</li> </ul>	Allow reduction <b>and</b> oxidation	2
9(c)	<ul style="list-style-type: none"> <li>• identification: green precipitate is iron(II) hydroxide (1)</li> </ul> Explanation: <ul style="list-style-type: none"> <li>• precipitate turns brown because iron(II) hydroxide is oxidised by oxygen (and water) in the air (1)</li> <li>• to form iron(III) hydroxide (1)</li> </ul>	Allow $\text{Fe}(\text{OH})_2$ / $\text{Fe}(\text{H}_2\text{O})_4(\text{OH})_2$  Allow $\text{Fe}(\text{OH})_3$ / $\text{Fe}(\text{H}_2\text{O})_3(\text{OH})_3$	3

(Total for Question 9 = 8 marks)

Question number	Answer	Additional guidance	Marks
<b>10(a)</b>	An explanation that makes reference to the following points: <ul style="list-style-type: none"> <li>• they get more stable down the group (1)</li> <li>• because the size of the cations increases/charge density of cations decreases (1)</li> <li>• and so carbonate ions are less polarised (1)</li> </ul>		<b>3</b>
<b>10(b)</b>	<ul style="list-style-type: none"> <li>• rearrangement of equation (1)</li> <li>• calculation of <math>\Delta H_{\text{solution}}</math> (1)</li> </ul>	<u>Example of calculation</u> $- 2493 + \Delta H_{\text{solution}} = -1920 + (-2 \times 364)$ $\Delta H_{\text{solution}} = -155 \text{ (kJ mol}^{-1}\text{)}$ <p>Correct sign must be given in final answer</p> <p>Correct answer and sign with no working scores 2 marks</p>	<b>2</b>
<b>10(c)(i)</b>	An explanation that makes reference to the following points: <ul style="list-style-type: none"> <li>• breaking the lattice is endothermic and the hydration of ions is exothermic (1)</li> <li>• (therefore the dissolving of magnesium sulphate is exothermic) because the enthalpy of hydration (of the ions) is greater in magnitude than the lattice energy (of <math>\text{MgSO}_4</math>) (1)</li> </ul>		<b>2</b>
<b>10(c)(ii)</b>	<ul style="list-style-type: none"> <li>• <math>\Delta G^\ominus = -87 - (298 \times -0.210)</math>  <math>= -24(.42) \text{ (kJ mol}^{-1}\text{)}</math> (1)</li> <li>• since <math>\Delta G</math> is negative the process/reaction is spontaneous/feasible (1)</li> </ul>		<b>2</b>

Question number	Answer	Additional guidance	Marks												
*10(d)	<p>This question assesses a student's ability to show a coherent and logically structured answer with linkages and fully-sustained reasoning.</p> <p>Marks are awarded for indicative content and for how the answer is structured and shows lines of reasoning.</p> <p>The following table shows how the marks should be awarded for indicative content.</p> <table border="1" data-bbox="317 422 651 716"> <thead> <tr> <th data-bbox="317 422 473 578">Number of indicative marking points seen in answer</th> <th data-bbox="473 422 651 578">Number of marks awarded for indicative marking points</th> </tr> </thead> <tbody> <tr> <td data-bbox="317 578 473 606">6</td> <td data-bbox="473 578 651 606">4</td> </tr> <tr> <td data-bbox="317 606 473 633">5-4</td> <td data-bbox="473 606 651 633">3</td> </tr> <tr> <td data-bbox="317 633 473 661">3-2</td> <td data-bbox="473 633 651 661">2</td> </tr> <tr> <td data-bbox="317 661 473 688">1</td> <td data-bbox="473 661 651 688">1</td> </tr> <tr> <td data-bbox="317 688 473 716">0</td> <td data-bbox="473 688 651 716">0</td> </tr> </tbody> </table>	Number of indicative marking points seen in answer	Number of marks awarded for indicative marking points	6	4	5-4	3	3-2	2	1	1	0	0	<p>Guidance on how the mark scheme should be applied:</p> <p>The mark for indicative content should be added to the mark for lines of reasoning. For example, an answer with five indicative marking points that is partially structured with some linkages and lines of reasoning scores 4 marks (3 marks for indicative content and 1 mark for partial structure and some linkages and lines of reasoning).</p> <p>If there are no linkages between points, the same five indicative marking points would yield an overall score of 3 marks (3 marks for indicative content and no marks for linkages).</p>	<b>6</b>
Number of indicative marking points seen in answer	Number of marks awarded for indicative marking points														
6	4														
5-4	3														
3-2	2														
1	1														
0	0														

Question number	Answer	Additional guidance	Marks								
<p><b>*10(d) cont.</b></p>	<p>The following table shows how the marks should be</p> <table border="1" data-bbox="317 169 951 537"> <tr> <td data-bbox="317 169 716 301"></td> <td data-bbox="716 169 951 301">Number of marks awarded for structure of answer and sustained line of reasoning</td> </tr> <tr> <td data-bbox="317 301 716 405">Answer shows a coherent and logical structure with linkages and fully sustained lines of reasoning demonstrated throughout.</td> <td data-bbox="716 301 951 405">2</td> </tr> <tr> <td data-bbox="317 405 716 485">Answer is partially structured with some linkages and lines of reasoning.</td> <td data-bbox="716 405 951 485">1</td> </tr> <tr> <td data-bbox="317 485 716 537">Answer has no linkages between points and is unstructured.</td> <td data-bbox="716 485 951 537">0</td> </tr> </table> <p>awarded for structure and lines of reasoning.</p> <p><b>Indicative content</b></p> $(\Delta G^{\circ}_{\text{solution}} = \Delta H^{\circ}_{\text{solution}} - T\Delta S^{\circ}_{\text{system}})$ <ul style="list-style-type: none"> <li>• for BaSO<sub>4</sub>: <math>\Delta H^{\circ}_{\text{solution}}</math> and <math>-T\Delta S^{\circ}_{\text{system}}</math> are both positive (1)</li> <li>• for CaSO<sub>4</sub>: <math>\Delta H^{\circ}_{\text{solution}}</math> is negative and <math>-T\Delta S^{\circ}_{\text{system}}</math> is positive (1)</li> <li>• but the magnitude of <math>-T\Delta S^{\circ}_{\text{system}}</math> is greater than that of <math>\Delta H^{\circ}_{\text{solution}}</math> (1)</li> <li>• therefore <math>\Delta G^{\circ}_{\text{solution}}</math> for both salts is positive (1)</li> <li>• when <math>\Delta G^{\circ}_{\text{solution}}</math> is positive the salt is only slightly soluble (1)</li> <li>• BaSO<sub>4</sub> is less soluble than CaSO<sub>4</sub> because <math>\Delta G^{\circ}_{\text{solution}}</math> is more positive (1)</li> </ul>		Number of marks awarded for structure of answer and sustained line of reasoning	Answer shows a coherent and logical structure with linkages and fully sustained lines of reasoning demonstrated throughout.	2	Answer is partially structured with some linkages and lines of reasoning.	1	Answer has no linkages between points and is unstructured.	0		
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Answer shows a coherent and logical structure with linkages and fully sustained lines of reasoning demonstrated throughout.	2										
Answer is partially structured with some linkages and lines of reasoning.	1										
Answer has no linkages between points and is unstructured.	0										

**(Total for Question 10 = 15 marks)**